Calculations Homework Name:

**The relative atomic mass (shortened to Ar) is a w\_\_\_\_\_\_\_\_\_ average of the mass of all the atoms of an element compared to carbon -12.**

1. If there are 2 atoms of magnesium-24 and 1 of magnesium-25.

What would the mean atomic mass of magnesium be then? \_\_\_\_

1. If there are 4 atoms of neon-20 and 1 of neon-21.

What would the mean atomic mass of magnesium be then? \_\_\_\_

**Calculate the formula masses of the following compounds:**

1. NaOH \_\_\_\_\_
2. H2CO3 \_\_\_\_\_
3. K3PO4 \_\_\_\_\_

**The mass of one mole is equal to the formula mass. Calculate the mass of one mole of:**

1. C6H6 \_\_\_\_\_
2. (NH4)2SO4 \_\_\_\_\_
3. Mg(OH)2 \_\_\_\_\_

**To calculate the following, multiply the formula mass by the number of moles:**

1. 2 moles of C3H6 \_\_\_\_\_ x 2 = \_\_\_\_\_\_\_\_\_
2. 0.5 moles of NH4NO3 \_\_\_\_\_ x \_ = \_\_\_\_\_\_\_\_\_
3. 0.1 moles of Mg(OH)2 \_\_\_\_\_ x \_ = \_\_\_\_\_\_\_\_\_

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