Reactivity Series Homework

Use the diagrams of the metals below reacting with dilute acid to put metals A, B, C, D & E in order of reactivity. [2]

 A B C D E

|  |  |
| --- | --- |
| Least Reactive | Most Reactive |
|  |  |  |  |  |

Imagine a new metal, given the symbol J, has been discovered. It lies between Lithium & Calcium in the reactivity series.

Describe the reaction of excess J with dilute Nitric Acid and give a word equation and balanced symbol equation (metal J forms 2+ ions)

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J is added to a solution of copper (II) sulphate.

Explain what you would expect to see happen, including a balanced symbol equation with state symbols.

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Write an ionic equation, including state symbols, showing what happens in the above reaction (HT)

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Explain which species is oxidised and which is reduced using ionic half equations (HT)

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